**VSEPR Online Activity**

What does VSEPR stand for?

The VSEPR theory tells us that areas of a molecule that have a higher concentration of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ will repel each other and that these repulsions will \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of the molecule.

On slide 6 when you move the model, why do the red and blue atom always stay on opposite sides of the white atom?

In a molecule with three peripheral atoms, each atom will take positions \_\_\_\_\_\_\_\_\_\_\_° apart.

Before moving on to slide 10, draw what you think the molecule will look like with 4 peripheral atoms.

When you have four peripheral atoms, what is the angle between the bonds? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Bond angles for CO2: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ shape: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Bond angles for NO3-: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ shape: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Bond angles for SO2: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ shape: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Bond angles for CH4: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ shape: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Bond angles for NH3: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ shape: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Bond angles for H2O: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ shape: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Now go back to each one and draw their perspective drawings below. Be sure to label your drawings: