Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Introduction to Stoichiometry Video

Stoichiometry allows us to compare quantities of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ together. So that if we have a certain amount of reactant we can figure out how much of a product was produced. Or if we know how much product was produced, we can figure out how much \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ must have been used up in order to produce that product.

We can’t easily count out \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ or \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. But what we can do is count out and take quantity of a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ number of them.

So we have to have a go-between between mass and number. And that of course is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

When you start with the mass of a particular reactant, you have to convert it to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. This is where our \_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ comes in. This allows us to get to the mass of the products.

Ultimately, stoichiometry allows us to take the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of a reactant (or product), use the \_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in order to figure out how many moles of the product (or reactant) we should produce, and of course if we are starting or ending with mass, it is the molar mass which allows us to do that.

If you know the formula of a compound you can calculate its \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

You’re going to use the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ from the balanced chemical equation to figure out the mass or volume of a product (or reactant).

Ultimately what this allows us to do is figure out how much \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ is going to be produced.