

## Molar Mass Worksheet

Calculate the molar mass of the following chemicals:

- 1)  $\text{Cl}_2$
- 2)  $\text{KOH}$
- 3)  $\text{BeCl}_2$
- 4)  $\text{FeCl}_3$
- 5)  $\text{BF}_3$
- 6)  $\text{CCl}_2\text{F}_2$
- 7)  $\text{Mg}(\text{OH})_2$
- 8)  $\text{UF}_6$
- 9)  $\text{SO}_2$
- 10)  $\text{H}_3\text{PO}_4$
- 11)  $(\text{NH}_4)_2\text{SO}_4$
- 12)  $\text{CH}_3\text{COOH}$
- 13)  $\text{Pb}(\text{NO}_3)_2$
- 14)  $\text{Ga}_2(\text{SO}_3)_3$

## Moles / Representative Units Conversions

Compound	Representative Unit
Copper II Chloride	
SO <sub>2</sub>	
Potassium	
Ca <sup>2+</sup>	

1. How many moles of Pb are in  $9.3 \times 10^{15}$  atoms of Pb?
2. How many molecules are in 2.5 mol CO<sub>2</sub>?
3. How many moles are in  $3.52 \times 10^{24}$  molecules of water?
4. How many atoms of zinc are in 0.60 mol of zinc?
5. How many moles of sodium chloride are in  $3.7 \times 10^{25}$  formula units of sodium chloride?