

Name: _____

Date: _____

Mole Ratio Worksheet

1) Given this equation: $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$, balance it and write the following molar ratios:

a) N_2 / H_2

b) N_2 / NH_3

c) H_2 / NH_3

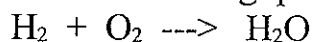
2) Given the following equation: $\text{H}_2 + \text{S}_8 \rightarrow \text{H}_2\text{S}$, balance it and write the following molar ratios:

a) $\text{H}_2 / \text{H}_2\text{S}$

b) H_2 / S_8

c) $\text{H}_2\text{S} / \text{S}_8$

3) Balance and answer the following questions for this equation:



a) What is the $\text{H}_2 / \text{H}_2\text{O}$ molar ratio?

b) Suppose you had 20 moles of H_2 , how many moles of H_2O could you make?

c) What is the $\text{O}_2 / \text{H}_2\text{O}$ molar ratio?

d) Suppose you had 20 moles of O_2 , how many moles of H_2O could you make?

4) Balance and use this equation: $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$, for the following problems

a) If you used 1 mole of N_2 , how many moles of NH_3 could be produced?

b) If 10 moles of NH_3 were produced, how many moles of N_2 would be required?

c) If 3.00 moles of H_2 were used, how many moles of NH_3 would be made?

d) If 0.600 moles of NH_3 were produced, how many moles of H_2 are required?