Covalent Bonding Online Activity

Answer the two questions on slide 3:

1.

2.

In a covalent bond, the atoms are not really sharing electrons as much as they are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ over them. This creates a situation in which both atoms are \_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

In a covalent bond, each of the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ atoms will have a strong attraction for the other atom’s electrons, but will also tend to hold onto its own electrons.

Which element is considered a nonmetal but is not grouped with all the other nonmetals?

Draw the graph of potential energy on slide 15:

Why does potential energy decrease as the atoms move closer to each other?

At what point does the potential energy suddenly increase (what causes this)?

The point at which the potential energy is the lowest and most stable can determine the \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_.

Lower in energy = \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_

A double bond represents a total of \_\_\_\_\_\_\_\_\_ shared electrons.

A triple bond represents a total of \_\_\_\_\_\_\_\_\_ shared electrons.

Naming compounds:

The second element always ends in \_\_\_\_\_\_\_\_\_\_\_\_

The prefix mono- is not used if there is only one atom of the \_\_\_\_\_\_\_\_\_\_\_\_ element.