

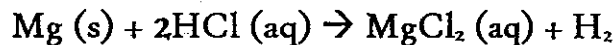
Name: \_\_\_\_\_

KEY

Hour: \_\_\_\_\_

## Chemical Equations

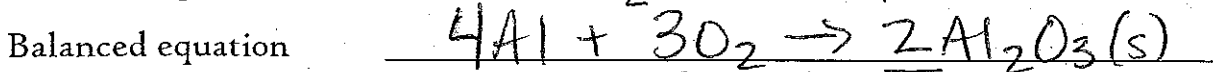
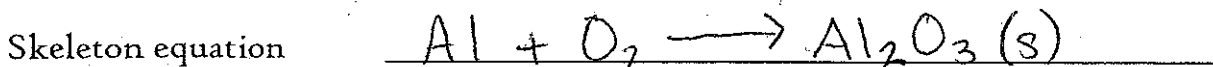
Explain in words what each of the symbols mean that are from the following chemical reaction:



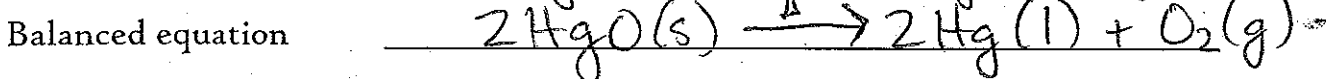
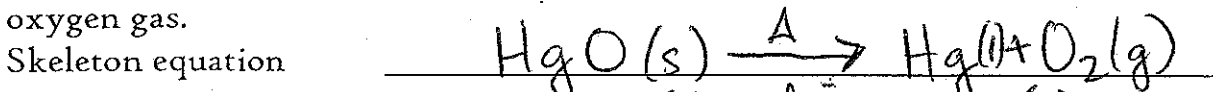
- |   |  |
|---|--|
| 1. Mg <u>Magnesium</u>                                      | 6. (aq) <u>dissolved in water</u>  |
| 2. (s) <u>in a solid state</u>                              | 7. $\rightarrow$ <u>yields, reacts to produce</u>                        |
| 3. HCl (aq) <u>hydrochloric acid in an aqueous solution</u> | 8. $\text{MgCl}_2$ (aq) <u>Magnesium chloride in an aqueous solution</u> |
| 4. 2 <u>there are 2 of the unit</u>                         | 9. $\text{H}_2$ <u>two hydrogen atoms</u>                                |
| 5. + <u>combined with, and</u>                              | 10. (g) <u>in a gaseous state</u>  |

Write a skeleton and then a balanced equation for the reactions below

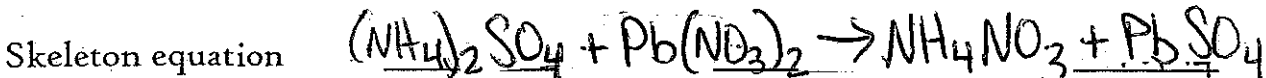
11. Aluminum metal burns in pure oxygen to produce solid aluminum oxide.



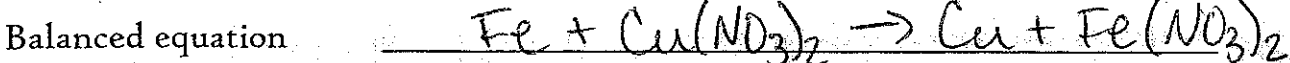
12. When solid mercury (II) oxide is heated, it breaks down to form liquid mercury and oxygen gas.



13. A solution of ammonium sulfate added to a solution of lead (II) nitrate, forms soluble ammonium nitrate and solid lead (II) sulfate.

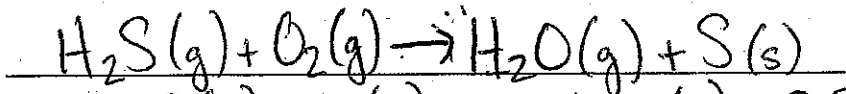


14. Copper metal and iron (II) nitrate in solution are formed when iron metal is added to a solution of copper (II) nitrate.

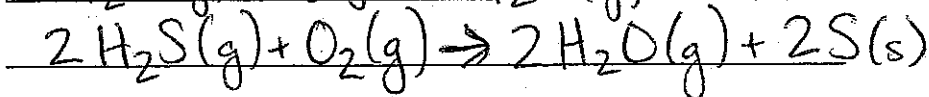


15. Hydrogen sulfide gas reacts with oxygen gas to form water vapor and solid sulfur.

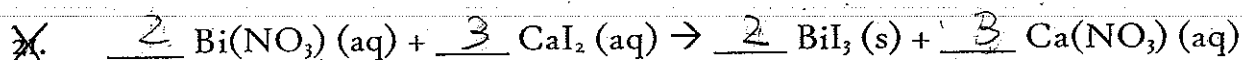
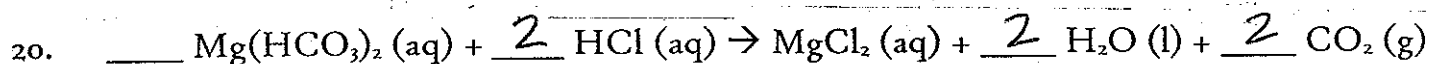
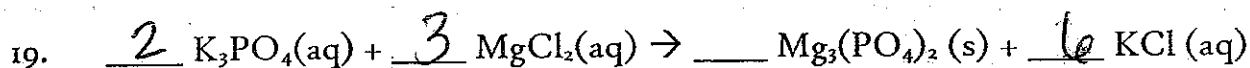
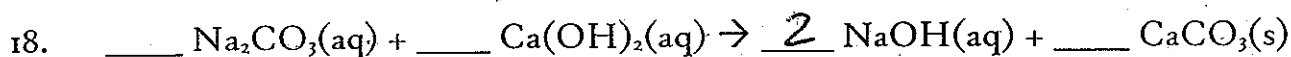
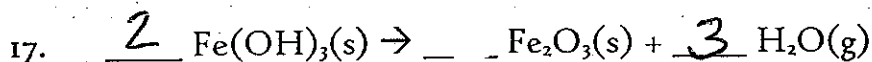
Skeleton equation



Balanced equation



Balance the following equations



Choose the correct symbol for the type of reaction. Place that answer in the blank at the beginning of each equation and then balance each equation correctly.

S = synthesis

SD = single displacement

D = decomposition

DR = double replacement

C = combustion

