Chapter 6 Study Guide

1. Be able to identify these groups on the periodic table:

Alkali metals Alkaline metals Transition metals Metalloids

Nonmetals Halogens Noble Gases

1. Be able to identify these trends on the periodic table using arrows:

Electronegativity ionization energy atomic radius

1. Define electronegativity in your own words:
	1. Explain WHY this trend changes from left to right and from top to bottom across the table:
2. Define ionization energy in your own words:
	1. Explain WHY this trend changes from left to right and from top to bottom across the table:
3. Define atomic radius in your own words:
	1. Explain WHY the atomic radius changes from left to right and from top to bottom across the table:
4. What is the difference between a cation and an anion?
	1. What is the difference in size between a cation and an atom of the same element?
	2. What is the difference in size between an anion and an atom of the same element?
5. Which direction do periods run on the periodic table? (left to right or top to bottom)
	1. What can these numbers tell you?
6. Which direction do groups run on the periodic table? (left to right or top to bottom)
	1. What changes as these numbers increase?