

13.2 The Nature of Liquids

According to kinetic theory there are no attractions between the particles of a gas, however there is between liquid particles. This is what gives liquids a definite volume.

Vaporization - the conversion of a liquid to a gas

Evaporation - the conversion of a liquid to a gas NOT
↓ due to boiling

is a cooling process. Liquids that become a gas do so because there was enough of an increase in kinetic energy to overcome their attractive forces. As the higher energy particles "escape" into the air the average temp of the remaining liquid is going to be lower.

Vapor pressure - a measure of the force exerted by a gas above a liquid.

Dynamic equilibrium - will occur in a system at constant vapor pressure with the same rate of evaporation of the liquid + condensation of a gas/vapor.

Boiling point (bp) - the temperature at which the vapor pressure of the liquid is just equal to the external pressure on the liquid

